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mostly empty space •

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gas are

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Heat

Thermodynamics

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2. Gas particles are in random, constant, straight line motion. 3. When particles collide with each other (or any container), the collisions are said to be elastic. 4. The volume that gas

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particles take up is negligible. And from the distance between particles is relatively great.

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ideal gas law has a
limited precision for
predicting the
properties of gases.

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BEHAVIOR OF
GASES REVIEW

Page 102 Chemistry

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Schools 11. Calculate

the new temperature

of a gas when 1500

mL at 25°C is

suddenly compressed

to 500 mL. Charles'

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Behavior Of

Law $K_m L K_m L V T V$

$T 100 (1500) (298$

$)(500) 1 1 2 2 12. A$

flask contains 34.6

kPa of CO_2

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Packet Answer Key

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Review 2 SECTION 2

BEHAVIOR OF

GASES 1. a measure

of how fast the

particles of an object

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Behavior Of

are moving 2. when it
is heated 3.

Temperature of gas
particles Energy of
gas particles Volume
of gas particles 1)
20° C Particles have
the smallest amount
of energy. Volume is
smallest. 2) 50° C
Particles have more

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Behavior Of Gases

Review 2 Answers

properties of gases.

The imprecision is known as the non-ideal behavior of gas,

and the van der

Waals equation. $(P +$

$n^2a/V^2)(V - nb) =$

nRT . has been

introduced to deal

with non-ideal

behavior of gases in

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Ideal gas law. Gases 2-

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Answers Page 8/26

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BEHAVIOR OF

GASES 1. a measure

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Behavior Of

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Answers
Temperature of gas
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20° C Particles have
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smallest. 2) 50° C
Particles have more

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GASES 1. a measure

of how fast the

particles of an object

are moving 2. when it

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is heated 3. Review 2

Answers
Temperature of gas particles Energy of gas particles Volume of gas particles 1) 20° C Particles have the smallest amount of energy. Volume is smallest. 2) 50° C Particles have

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20° C Particles have
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Particles have more
energy

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utterly discover a new
experience and
endowment by
spending more cash.
still when? complete
you acknowledge that
you require to get
those every needs
subsequent to having

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feel it, but air has pressure. The gases in Earth's

atmosphere exert pressure against everything they

contact. The

atmosphere rises high above Earth's

surface. It contains a huge number of individual gas

particles. As a result, the pressure of the

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tower of air above a 2
given spot on Earth's
surface is substantial.

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