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AP Chemistry Chapter 14 Answers – Zumdahl 14.41 pH = -log[H +]; pOH = -log[OH-]. At 25 oC, pH + pOH = 14.00; For Exercise 13.37: a. pH = -log[H +] = -log(1.0 x 10-7) = 7.00; pOH = 14.00 – pH = 14.00 – 7.00 = 7.00 b. pH = -log(8.3 x 10-16) = 15.08; pOH = 14.00 – 15.08 = -1.08 c. pH = -log(12) = -1.08; pOH = 14.08 – (-1.08) = 15.08 d.

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AP Chemistry Chapter 14, Chemical Kinetics - 1 - Change in time Change in the concentration of B Average rate with respect to B = Chapter 14, Chemical Kinetics 14.1 Factors that Affect Reaction Rates • Chemical kinetics = the study of how fast chemical reactions occur. • Factors which affect rates of reactions:

Chapter 14, Chemical Kinetics

AP CHEMISTRY – MCHS Page | 9 14.5 Temperature and Rate (EOCC ' S 47 - 59 page 620) • Most reactions speed up as temperature increases. • We can illustrate this with chemiluminescent Cyalume® light sticks. • A chemiluminescent reaction produces light. • Two light sticks are placed in water, one at room temperature and one in ice.

Chapter 14, Chemical Kinetics

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Answer: Using Equation 14.14, we find [NO 2] = 1.00 x 10 –3 M Practice Exercise Solution (continued) 18. Sample Exercise 14.9 Determining the Half-Life of a First-Order Reaction The reaction of C 4 H 9 Cl with water is a first-order reaction.

AP Chemistry Chapter 14 Sample Exercises

Answer: carbon, by 18.1 ° C Calculate the work done for the expansion of oxygen gas from 2.5 liters to 3.7 liters against a pressure of 720 torr. Answer: -1.14 liter-atm 20.0 mL of pure water at 285 K is mixed with 48 mL of water at 315 K. What is the final temperature of the mixture in kelvins?

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